Academic Year العام الدراسي	**2025/2026**
Term الفصل	**1**
Subject المادة	**Chemistry /Inspire** الكيمياء / انسباير
Grade الصف	**10**
Stream المسار	**Advanced** المتقدم
Number of MCQ عدد الأسئلة الموضوعية	**30**
Marks of MCQ درجة الأسئلة الموضوعية	**From 2- 4**
Number of FRQ عدد الأسئلة المقالية	**0**
Marks per FRQ الدرجات للأسئلة المقالية	**0**
Type of All Questions نوع كافة الأسئلة	**MCQ/** الأسئلة الموضوعية
Maximum Overall Grade الدرجة القصوى الممكنة	**100**
Exam Duration - مدة الامتحان	**120 minutes**
Mode of Implementation -طريقة التطبيق	**SwiftAssess**
Calculator الآلة الحاسبة	**Allowed** مسموحة

		Question* السؤال*	**Learning Outcome/Performance Criteria**** نتائج التعلم/ معايير الأداء**	**Reference(s) in the Student Book (English Version& Arabic Version)** المرجع في كتاب الطالب (النسخة الإنجليزية والنسخة العربية)	
				Example/Exercise مثال/تمرين	**Page** الصفحة
MCQ - الأسئلة الموضوعية	1	CHM.5.1.01.001.04	Describe Bohr's model of the hydrogen atom	Textbook + Figure 11	13, 14
	2	CHM.5.1.01.001.07	Differentiate between the ground and excited states of an atom	Textbook + Figure 11	14
	3	CHM.5.1.01.001.07	Differentiate between the ground and excited states of an atom	Textbook + Figure 10	13
	4	CHM.5.1.01.002.01	Describe the four quantum numbers and the significance of each to help track of electrons present in an atom	Textbook + Figure 16	20
	5	CHM.5.1.01.003.02	Explain the Aufbau principle, the Pauli exclusion principle, and Hund's rule	Textbook+Figure 18	23
	6	CHM.5.1.01.003.02	Explain the Aufbau principle, the Pauli exclusion principle, and Hund's rule	Textbook	24
	7	CHM.5.1.01.003.05	Write the electron configuration, orbital diagram, and noble gas notation of different elements	Textbook + Tables 4 and 5	25, 26
	8	CHM.5.1.02.001.02	Determine the number of valence electrons in atoms	Textbook + Table 6 + Example problem 3+ Practice problems	28, 29
	9	CHM.5.1.01.004.04	Identify the information that can be displayed in the periodic table	Textbook + Figure 4	39, 40, 41
	10	CHM.5.1.01.010.01	Describe the general properties of metals, non-metals and metalloids and their uses	Textbook+ Figures 4, 5, 6	40
	11	CHM.5.1.01.008.02	Use the electron configuration notation, orbital notation, and noble gas notation of an element (Z 1 – 36) to identify the location of an element in the periodic table (period, group and block)	Textbook+ Example problem 1	46
	12	CHM.5.1.01.004.10	Use the electron configuration notation, orbital notation, and noble gas notation of an element (Z 1 – 36) to identify the name and symbol of an element	Textbook+ Tables 3 and 4 + Figures 8 and 9	43, 44, 45
	13	CHM.5.1.01.004.12	Conclude the electron configuration of elements based on their position in the periodic table (period and group)	Textbook+ Example problem 1	46
	14	CHM.5.1.01.009.14	Describe the trend of ionic radii across a period and down a group of the periodic table	Textbook + Figures 11, 12, 13 + Example problem 2 + Practice problems	49, 50, 51
	15	CHM.5.1.01.009.10	Explain the trend of first ionization energy across a period (Exceptions between Groups 2 & 3 and 5 & 6 are included), and down a group of the periodic table	Textbook + Figure 17	54
	16	CHM.5.1.01.011.02	Explain the periodic trend of electronegativity across a period and down a group of the periodic table	Textbook + Figure 19	56, 57
	17	CHM.5.1.02.022.04	Relate the charge of an ion to the group number of an element in the periodic table	Textbook + Tables 1 and 2 + Figure 2	63, 64, 65
	18	CHM.5.1.02.022.03	Describe how ions (cations and anions) form to fulfill the octet rule	Textbook + Tables 1 and 2 + Figure 2	63, 64, 65
	19	CHM.5.1.02.022.07	Use the Lewis diagram (electron–dot structure) to explain how elements from the periodic groups combine to form an ionic compound	Text book + Table 4 + Practice problems	68, 69, 70
	20	CHM.5.1.02.022.08	Explain the physical properties of ionic compounds as melting point and boiling point, conductivity when solid, molten or aqueous, and its solubility in water	Textbook + Table 5	72, 73
	21	CHM.5.1.02.022.09	Explain the effect of applying an external force on the crystal	Textbook +Figure 8	72, 73
	22	CHM.5.1.01.013.02	Write the chemical formula (using the stock method) of an ionic compound containing monoatomic and polyatomic ions	Textbook + Tables 7, 8 + Example problems 1, 2	75, 76, 77
	23	CHM.5.1.01.013.02	Write the chemical formula (using the stock method) of an ionic compound containing monoatomic and polyatomic ions	Textbook + Tables 9, 10 + Example problem 3 + Practice problems	78, 79
	24	CHM.5.1.01.013.01	Write the chemical name of an ionic compound containing monoatomic and polyatomic ions (including oxyanions)	Text book+ Tables 10, 11 + Practice problems	79, 80
	25	CHM.5.1.02.023.02	Contrast the structures of ionic and metallic compounds	Textbook + Figure 10	82, 83
	26	CHM.5.1.02.023.03	Explain some physical properties of metals (Melting and boiling points, Thermal and electrical conductivity, Malleability, ductility,	Textbook + Figure 11 + Table 12	83, 84
	27	CHM.5.1.01.011.08	Identify type of elements involved in the covalent bond with the movement of electrons	Textbook+ Figure 5 + Example problem 1+ Practice problems	94, 95
	28	CHM.5.1.02.007.04	Identify the relationship between the type of a covalent bond (single, double and triple) and its bond length, bond strength and b	Textbook + Figure 10 + Table 1	98
	29	CHM.5.1.01.014.02	Name a binary molecular compound based on its molecular formula (up to deca–)	Textbook + Table 2 + Example problem 2 + Practice problems	100, 101
	30	CHM.5.1.02.002.02	Draw Lewis structures for a number of polyatomic ions	Textbook + Example problems 3, 4, 5 + Figure 14 + Practice problems	109, 110, 111

*	**Questions might appear in a different order in the actual exam**
*	قد تظهر الأسئلة بترتيب مختلف في الامتحان الفعلي
**	**As it appears in the textbook(UAE Edition Grade 10 Advance-Inspire Chemistry Student Edition) , LMS, and (Main_IP).**
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